

Reviews in Mineralogy and Geochemistry



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Title, headings, and subheadings. Please use the following style.

Tables Prepare each table in Microsoft Word or Excel with its proper (and informative) heading in place (see example below). Do not use font sizes smaller than 11 point.

Pyroxenes

Title

James. A. Author and Carol C. Coauthor

Author List

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Email addresses

ROCK-FORMING PYROXENES

1st order

Assemblages of natural pyroxenes [no period, no caps]

2nd order

Subsolidius assemblages. (An indented, italicized, bold heading is the form for a third-order heading and that the text should follow after it exactly as this text does.) 3rd order

Table 23. The composition of hot chocolate.

[center heading, using 11-pt font]

Constituent	Weight %
Milk	XX
Cocoa	уу
Sugar	ZZ

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$$\sum t_1 z = 6.589 [\{T_{1z} - O\} - 1.605]$$
(32)

Use the following style in both the equation and surrounding text:

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SCALAR VARIABLES are italicized

VECTORS and MATRICES are bold faced

Abbreviations and symbols should be as below. There is a space between all numbers and units:

```
Common abbreviations (note Roman type)
                                                                         Density (D or ρ)
cf. = compare
                                                                         g/cm^3 = gram(s) per cubic centimeter
e.g. = for example
et al. = and others
                                                                         Concentration (use square brackets [ ] around chemical
etc. = and other things
                                                                            symbol, e.g., [Ca<sup>-2</sup>])
i.e. = that is
                                                                         % = percent [no space prior]
p. = page(s)
                                                                         % = per thousand [no space prior]
vs. = versus
                                                                         wt% = weight percent
                                                                         mol% = mole percent
Length (1)
                                                                         vol% = volume percent
m = meter(s)
                                                                         M = \text{molar concentration (mol/L)}
cm = centimeter(s) (1 cm = 10^{-2} m)
                                                                         m = \text{molal concentration (mol/kg)}
mm = millimeter(s) (1 mm = 10^{-3} m)
                                                                         ppm = parts per million
\mu m = micrometer(s) (1 \mu m = 10^{-6} m) NOT micron or \mu
                                                                         ppb = parts per billion
nm = nanometer(s) (1 nm = 10^{-9} m)
Å = angstrom(s) (1 Å = 10^{-10} m)
                                                                         Time (t)
pm = picometer(s) (1 pm = 10^{-12} m)
                                                                         s = second(s)
                                                                         min = minute(s)
Mass (m)
                                                                         h = hour(s)
g = gram(s)
                                                                         d = day(s)
kg = kilogram(s)
                                                                         yr = year(s)
mg = milligram(s)
                                                                         Ma = million years
                                                                         Ga = billion years (ago or date)
\mu g = microgram(s)
                                                                         m.y. = million years
                                                                         b.y. = billion years (duration)
Volume (V)
L = liter(s)
mL = milliliter(s)
                                                                         Frequency (f or v)
cm^3 = cubic centimeter(s)
                                                                         Hz = hertz
Å^3 = cubic angstrom(s)
                                                                         kHz = kilohertz
                                                                         MHz = megahertz
Pressure (P)
Pa = pascal(s)
                                                                         Miller indices--note that letter symbols are italicized but
kPa = kilopascal(s)
                                                                            numeric values are not, e.g., [hkl] but [100]
MPa = megapascal(s)
                                                                         hkl = diffraction symbol
GPa = gigapascal(s)
                                                                         (hkl) = face symbol
bar or bars (no abbreviation)
                                                                         \{hkl\} = form symbol
kbar = kilobar(s)
                                                                         [hkl] = edge or zone symbol
atm = atmosphere(s)
                                                                         Unit-cell measurements:
Temperature (T)
                                                                         a, b, c = edge lengths
K = Kelvin(s)
                                                                         \alpha, \beta, \gamma = angles
°C = degrees Celsius [space before degree sign. c.f. angles
                                                                         \mathbf{a}, \mathbf{b}, \mathbf{c} = \text{vectors--note boldface}
where there is no space between digit and degree sign]
                                                                         Optical measurements:
Energy:
                                                                         dispersion (r, v),
J = joule(s)
                                                                         optic axial angle (2V, 2V_x, 2V_z);
kJ = kilojoule(s) (preferred)
                                                                         principal vibration directions (E, O; X, Y, Z);
cal = calorie(s)
                                                                         refractive indices (n; \varepsilon, \omega; \alpha, \beta, \gamma)
kcal = kilocalorie(s)
```

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Book:

Tossell JA, Vaughn DJ (1992) Theoretical Geochemistry: Application of Quantum Mechanics in the Earth and Mineral Sciences. Oxford University Press, New York

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Keil K (1969) Meteorite composition. *In*: Handbook of Geochemistry. Vol 1. Wedepohl KH (ed) Springer, Berlin, p 78-115

Edited Book:

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Conference Proceedings or Abstract:

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FIGURES

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