

Phosphorus speciation changes across the glass transition in highly polymerized alkali silicate glasses and melts

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ABSTRACT

In situ, high-temperature Raman spectroscopy of $\text{Na}_2\text{O}\cdot 9\text{SiO}_2$ glasses and melts with and without 2 mol% P_2O_5 has been conducted to 1491 K (near the liquidus temperature). In the temperature range from 298 to ~ 1000 K (below the glass-transition temperature range), orthophosphate (PO_4^{3-}) complexes appear to be the dominant phosphate species present. As temperature increases above ~ 1000 K, the orthophosphate complex is partially replaced by pyrophosphate complexes ($\text{P}_2\text{O}_7^{4-}$) so that at ~ 1200 K $>70\%$ of the P exists in the latter form. The proportion of pyrophosphate species does not change significantly as the temperature further increases. This change in phosphate speciation is accompanied by an increase in the Q^3/Q^4 silicate species abundance ratio. At the highest temperatures there is also evidence for Q^2 species in the melts.

The transformation from orthophosphate to pyrophosphate implies depolymerization of the silicate network because the P/O ratio in the $\text{P}_2\text{O}_7^{4-}$ complex is smaller than in PO_4^{3-} . This depolymerization is consistent with the change in Q^3/Q^4 ratio in the silicate. Therefore, use of P speciation data from glasses would lead to overestimation of the degree of polymerization of P-bearing silicate melts.

INTRODUCTION

The geochemical behavior of P in natural magmatic systems is receiving increasing attention (e.g., Harrison and Watson 1984; Pichavant et al. 1992; Toplis et al. 1994; Wolf and London 1994). Although generally present in comparatively low concentrations in natural magmas ($<2\text{--}3$ wt% P_2O_5), P strongly modifies both the physical and chemical properties of silicate liquids (e.g., Wyllie and Tuttle 1964; Kushiro 1975; Watson 1976; Ryerson and Hess 1980; London 1987; Toplis et al. 1994).

Results of in situ examination of silicate melt and glass structures at high temperature indicate that above the glass transition there is a gradual change in the proportion of the silicate structural units (Q species) with increasing temperature. The enthalpy of the relevant speciation reactions typically is on the order of 10–40 kJ/mol (e.g., Stebbins et al. 1992; McMillan et al. 1992; Mysen and Frantz 1992, 1993). These studies also indicate that the structure of silicate glasses reflects the structure frozen in at the glass-transition temperature. Glass structure cannot be used, therefore, for quantitative description of silicate liquids. To characterize the structural role of P in silicate melts at high temperature, in situ, high-temperature studies similar to those conducted for silicate speciation in melts is needed. This is the purpose of the present study.

Results from structural studies of glass indicate poten-

tially complex interactions between P, alkali and alkaline earths, Al, Si, and Fe (e.g., Mysen et al. 1981; Nelson and Tallant 1984; Dupree et al. 1988; Gan and Hess 1992; Mysen 1992). In light of this complexity, a database for the structural behavior of P in silicate melts at high temperature is needed. Choosing a simple silica-rich composition ($\text{Na}_2\text{O}\cdot 9\text{SiO}_2$) with fixed P content (~ 2 mol% P_2O_5) was the first step in this process. This composition provides a simple model for the role of alkali-silicate interaction in a melt that has a degree of polymerization ($\text{NBO}/\text{Si} = 0.1$) that is similar to that found in felsic magmatic liquids (granite and rhyolite).

EXPERIMENTAL METHODS

The starting material was prepared (denoted NS9) by heating a thoroughly mixed $\text{Na}_2\text{CO}_3 + \text{SiO}_2$ oxide to 1600 °C for about 1 h. The melt was then quenched to a glass. P was added as ammonium dihydrogen phosphate ($\text{NH}_4\text{H}_2\text{PO}_4$) in a proportion appropriate to yield NS9 + 2 mol% P_2O_5 in the final glass starting material. The glass + $\text{NH}_4\text{H}_2\text{PO}_4$ mixture was ground for 1 h under alcohol and transformed to an NH_4^+ , H_2O^- , PO_4^- -containing glass by subjecting this mixture, contained in sealed platinum containers, to 1700 °C at 10 kbar for 1 h in a solid-media, high-pressure apparatus. This procedure was employed to avoid volatilization of the P which would occur if the glass + $\text{NH}_4\text{H}_2\text{PO}_4$ mixture was subjected to the required temperatures in open sample containers. The temperature was then decreased to 1550 °C at 1 atm for ~ 30 min to remove the ammonium and H_2O from the high-pressure

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TABLE 1. Composition of starting materials (wt%)

Oxide	NS9	NS9 + 2 mol% P ₂ O ₅
SiO ₂	90.24	88.44
Al ₂ O ₃	0.02	0.21
Na ₂ O	9.42	9.05
P ₂ O ₅	—	3.95
Total	99.68	101.65

glass. The glass was checked for homogeneity and analyzed with a JEOL 8900 electron microprobe (Table 1). Also shown is the analyzed composition of the P-free base glass.

About 1 mg glass chips from the starting material were used for the high-temperature, in situ structural work, which was conducted with confocal, microRaman spectroscopy using the heating system described by Mysen and Frantz (1993). The sample was optically examined continuously during the high-temperature experiments with the petrographic microscope attached to the Raman spectrometer (625× magnification). This was done to ascertain that there was no bubble formation, phase separation, or crystal nucleation while the high-temperature Raman spectra were acquired.

About 500 mW of the 514 nm line from an Ar⁺ ion argon laser was used for sample excitation. Raman spectra were recorded with an EG&G Model 4000 CCD with acquisition times ranging from 30 to 60 s per CCD window. Two windows were needed to collect the spectra in the frequency range of first-order Raman scattering from these samples. The experiments were conducted by first recording a glass spectrum at 25 °C followed by the spectrum at the highest temperatures and then recording the

spectra at successively lower temperatures and finishing with another spectrum at 25 °C. The process was then reversed by heating the same sample with the same temperature steps until the highest temperature was reached.

The high-frequency portion of the spectra (700–1300 cm⁻¹) was curve fitted with the method described by Mysen (1992, 1996) to determine the appropriate number of lines, the line shape, and the line parameters for each band. All line parameters and the fraction of Gaussian component were determined by minimizing the χ^2 and maximizing the residual distribution using the algorithm by Davidson (1966).

The glass cannot be physically removed from the heater after chemical analysis without destruction of the heater. Instead, to evaluate whether compositional degradation of the sample occurred during the high-temperature experimental series, the initial and final glass spectra were compared. By calibrating the changes in Raman spectra of glasses along the join Na₂O-SiO₂, Mysen (1995) found that spectral changes reflecting ± 1 wt% change in Na₂O content could be detected with this method.

RESULTS AND DISCUSSION

The 298 K spectrum of NS9 glass (Fig. 1A) is similar to that reported by Mysen et al. (1982). With increasing temperature, most spectral features broaden, as do those of other alkali silicate melts at similar temperatures.

By dissolving ~2 mol% P₂O₅ in the glass, a strong, narrow peak near 940 cm⁻¹ dominates the spectrum (Fig. 1B). In comparison with the P-free glass (Fig. 1A), other spectroscopic changes are subtle. With increasing temperature the band near 940 cm⁻¹ becomes less intense,

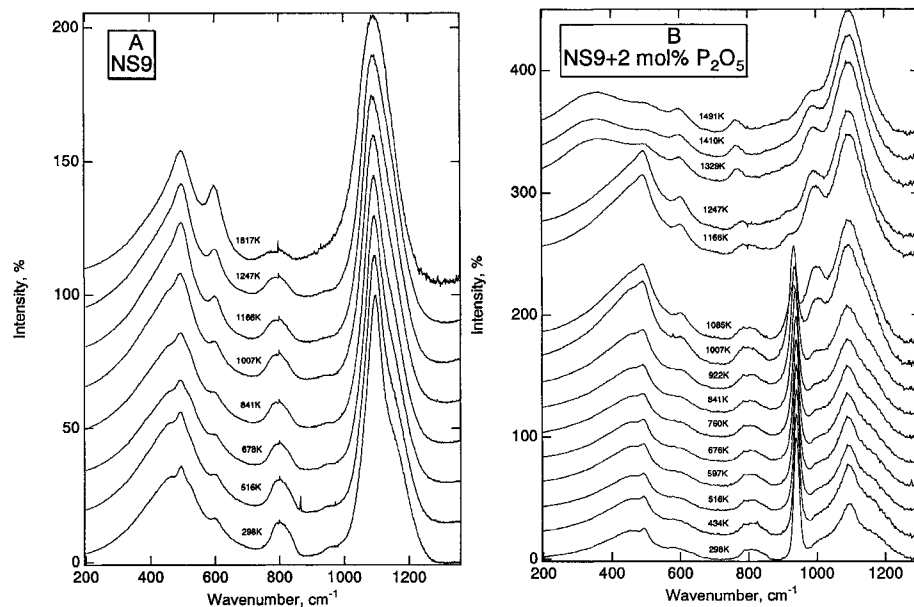


FIGURE 1. Unpolarized Raman spectra as a function of temperature (as indicated on individual spectra) for NS9 composition and NS9 + 2 mol% P₂O₅.

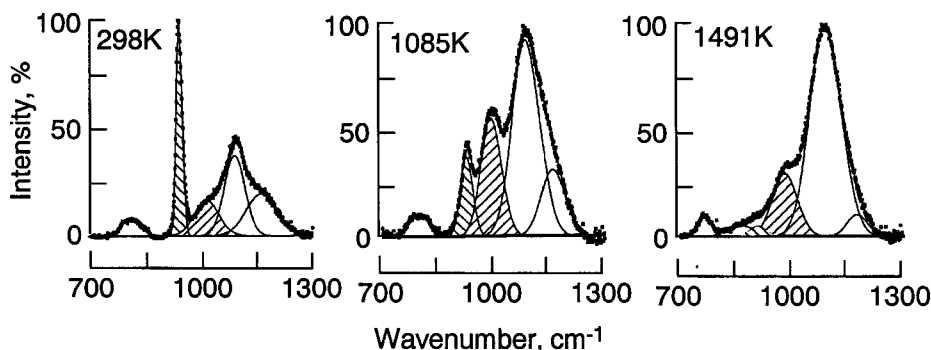


FIGURE 2. Examples of curve-fitted Raman spectra in the 700–1300 cm^{-1} frequency range at temperatures indicated. Hatchured bands are those assigned to P-O stretching in PO_3^{2-} and $\text{P}_2\text{O}_7^{4-}$ complexes.

and a new band grows near 1000 cm^{-1} . This effect becomes particularly evident as the temperature exceeds 900 K. The 940 cm^{-1} band decays to a barely visible shoulder at $T > 1247$ K, whereas the band near 1000 cm^{-1} remains strong. Above 1247 K additional spectral changes occur in the low-frequency region, with the peak centered near 500 cm^{-1} in the lower temperature spectra becoming much broader and less distinct, coupled with the development of a broad intensity maximum centered near 400 cm^{-1} (Fig. 1B). The broad band slightly above 800 cm^{-1} becomes sharper and is shifted to about 780 cm^{-1} . These spectroscopic changes are reversible across the temperature range examined.

To obtain more detailed information, the frequency range between 700 and 1300 cm^{-1} was fitted to five or six Gaussian lines (six lines were required for spectra taken at $T > 1200$ K). This frequency range includes those bands that can be assigned to Si-O⁻ and P-O⁻ stretching vibrations in species with different degrees of polymerization (different NBO/Si and NBO/P ratios). Three examples (298 K glass, a supercooled liquid at 1085 K, and liquid at 1491 K) are shown in Figure 2. The strongest band, near 1100 cm^{-1} , is assigned to Si-O⁻ stretching in Q³ structural units, and that near 1150 cm^{-1} to Si-O⁰ stretching in Q⁴ units. The weak band near 870 cm^{-1} in spectra taken at $T > 1200$ K is tentatively assigned to Si-O⁻ stretching in silicate species less polymerized than Q³. This species is here referred to as Q², but this assignment remains uncertain. The broad band near 800 cm^{-1} in the low-temperature spectra is present whether or not P is present (Fig. 1). The band is assigned to an in-cage O motion in fully polymerized structural units, Q⁴ (e.g., McMillan et al. 1992; Mysen 1995). The bands near 940 and 1000 cm^{-1} are assigned to P-O⁻ stretching in ortho- (PO_3^{2-}) and pyrophosphate ($\text{P}_2\text{O}_7^{4-}$), respectively (shaded in Fig. 2), consistent with ³¹P NMR and Raman spectroscopy of other P-bearing alkali and alkaline earth silicate glasses (Nelson and Tallant 1984; Dupree et al. 1988; Gan and Hess 1992; Mysen 1992). In silicate glasses in which the pyrophosphate complexes dominate, the presence of $\text{P}_2\text{O}_7^{4-}$ is also associated with increased Raman intensity near 400 cm^{-1} and a distinct

peak near 780 cm^{-1} (Nelson and Tallant 1984). Those spectroscopic features (broad 400 cm^{-1} band and sharp 780 cm^{-1} band) resemble the present spectra in the high-temperature range where the 1000 cm^{-1} band dominates over that near 940 cm^{-1} . It is concluded, therefore, that the spectroscopic changes at high temperature are due to vibrations associated with the greatly increased abundance of $\text{P}_2\text{O}_7^{4-}$ species in the melts in comparison with the lower temperature glasses.

The integrated intensity ratio (area) of the 940 cm^{-1} (PO_3^{2-}) and 1000 cm^{-1} ($\text{P}_2\text{O}_7^{4-}$) bands and of the 1100 cm^{-1} (Q³) and 1150 cm^{-1} (Q⁴) bands as a function of temperature is shown in Figure 3. At temperatures below ~1000 K, the area ratio, A_{940}/A_{1000} , may show a gradual decrease, although within the error of the data this decrease might be questionable. Between ~1000 and 1200 K this ratio decreases by >85%, followed by a very slow decrease at even higher temperature (Fig. 3A). The A_{1100}/A_{1150} ratio displays the opposite temperature dependence with a rapid increase in the 1000–1200 K temperature interval (Fig. 3B).

The liquidus temperature for the composition NS9 + 2 mol% P_2O_5 is about 1470 K [on the basis of the liquidus temperature of pure NS9 (Kracek 1930) adjusted for solution of 2 mol% P_2O_5 under the assumption of ideal mixing in the melt]. From the empirical relationship, $T_g = \frac{2}{3}T_{\text{liq}}$, this liquidus temperature, T_{liq} , yields a temperature for the onset of the glass transition, T_g , of 980 K. It would appear, therefore, that the rapid change in the integrated intensity ratios, A_{940}/A_{1000} and A_{1100}/A_{1150} , corresponds to the

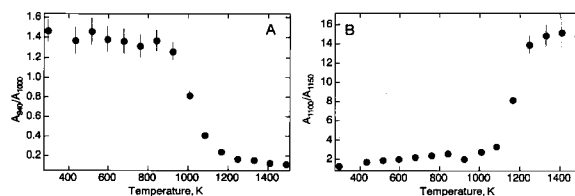


FIGURE 3. Area ratio of the 940 to 1000 cm^{-1} bands, A_{940}/A_{1000} (A), and of the 1100 and 1150 cm^{-1} bands, A_{1100}/A_{1150} (B), as a function of temperature.

temperature at which NS9 + 2 mol% P₂O₅ is transformed from a glass containing P mostly in the form of orthophosphate species, PO₄³⁻, to a supercooled liquid in which P occurs primarily as pyrophosphate units, P₂O₇⁴⁻. By assuming that the A₉₄₀/A₁₁₀₀ ratio is proportional to the concentration ratio of the two phosphate species in the glass and melt, the mole fraction of P₂O₇⁴⁻ increases from ~0.4 to ~0.9 across the glass-transition temperature range. This assumption is justified because the 940 and 1000 cm⁻¹ bands both are assigned to P-O stretching in highly depolymerized phosphate species. It is expected, therefore, that their Raman cross sections are similar, resulting in proportionality between area ratio and mole-fraction abundance ratio. The 1100 and 1150 cm⁻¹ bands, in contrast, are assigned to different vibrations in species that are structurally quite different (Si-O⁻ and Si-O⁰ stretching in Q³ and Q⁴ species, respectively). The Raman cross sections for those vibrations differ by a factor of 6–10 (Mysen 1995). Consequently, although A₁₁₀₀/A₁₁₅₀ is positively correlated with the Q³/Q⁴ abundance ratio, the two ratios are not proportional.

This change of phosphate complexing across the glass-transition temperature range requires interaction with the silicate solvent. Although there may be a small fraction of Q² species present, the two principal species are of the Q³ and Q⁴ types, as evidenced in particular by the strong bands near 1100 and 1150 cm⁻¹, respectively. The integrated intensity ratio of these two bands, A₁₁₀₀/A₁₁₅₀, increases across the glass-transition temperature range (Fig. 3B), which suggests that the change in phosphate speciation is also associated with a change in silicate polymerization. By using stoichiometric notations, Si₂O₅²⁻ and SiO₂, for the Q³ and Q⁴ species, a simple expression describes the equilibrium involving silicate and phosphate species: P₂O₇⁴⁻ + Si₂O₅²⁻ ↔ 2PO₄³⁻ + 2SiO₂. It follows from these observations that as the phosphate species undergo polymerization, O is exchanged between the phosphate and silicate species, so that the silicate melts actually become less polymerized as the phosphate polymerizes.

It is clear from the above considerations that when discussing phosphate-solution mechanisms in silicate melts, the quantitative evaluation of the interaction between silicate and phosphate requires determination of the phosphate speciation in the melt at relevant temperatures. By employing speciation data from glasses for this purpose, the effect of dissolved P on the polymerization of a silicate melt leads to an overestimation of the change in degree of polymerization.

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