Gailhanou et al. (2013), Standard enthalpies of formation at 298.15 K of the constituents used for solution calorimetry experiments. – Appendix 1.

Table A-1. Standard enthalpies of formation $\Delta H_f°$ of the constituents, at 298.15 K.

Constituents	$\Delta H_f°$ kJ/mol	Reference
Gibbsite $Al(OH)_{3(cr)}$	-1293.13 (± 1.2)	[1], [2]
α-Quartz $SiO_{2(cr)}$	-910.7 (± 1.7)	[1], [3]
α-Cristobalite $SiO_{2(cr)}$	-906.03 (± 2.1)*	[1], [4]
Goethite $FeOOH_{(cr)}$	-560.7 (± 1.2)	[1], [5]
Ferrihydrite $FeOOH·0.225H_2O$	-606.4* (± 1.2)	[6]
$FeO_{(cr)}$	-272.04 (± 2.1)	[1], [7]
Portlandite $Ca(OH)_{2(cr)}$	-984.55 (± 1.3)	[1], [8]
Brucite $Mg(OH)_{2(cr)}$	-924.14 (± 0.4)	[1], [9]
$Na_2SiO_{3(cr)}$	-1561.43 (± 4.2)	[7]
$KOH_{(cr)}$	-424.7 (± 0.6)	[10]
$KNO_{3(cr)}$	-494.5 (± 0.4)	[10]
MnO	-385.2 (± 0.5)	[1], [11]
Rutile $TiO_{2(cr)}$	-944.0 (± 0.8)	[1], [12]

Note. * $\Delta H_f°$ value for $FeOOH·0.225H_2O$ was calculated according to the linear function given by Majzlan et al. (2004), $\Delta H_f^0 = -541.2 - 289.2\,x$, where x is the compositional variable in $FeOOH·x\,H_2O$ ($0.2 < x < 1$).

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