

Synthetic biotite oxidation under hydrothermal conditions

CAROLYN ROSE REBBERT*

Department of Geological Sciences, University of Oregon, Eugene, Oregon 97403-1272, U.S.A.

ELIZABETH PARTIN

Route 1, Burnley Station, Barboursville, Virginia 22923, U.S.A.

DAVID A. HEWITT

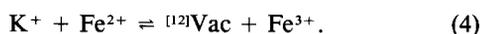
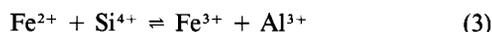
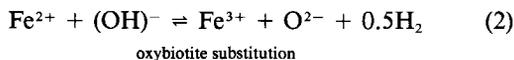
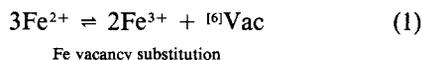
Department of Geological Sciences, Virginia Polytechnic Institute and State University, Blacksburg, Virginia 24061, U.S.A.

ABSTRACT

The oxidation of synthetic ferruginous biotite in hydrothermal systems was studied for two ideal compositions, annite, $\text{KFe}_3\text{AlSi}_3\text{O}_{10}(\text{OH})_2$, and siderophyllite, $\text{KFe}_{2.5}\text{Al}_2\text{Si}_{2.5}\text{O}_{10}(\text{OH})_2$. These biotite samples were annealed in Shaw-type vessels under controlled f_{H_2} at a total pressure of 1 kbar. The ratio of Fe^{2+} to total Fe, determined by wet-chemical analysis, is directly proportional to the f_{H_2} , regardless of temperature variations. For f_{H_2} values of < 1 bar to 102 bars, this ratio varies from 0.72 to 0.90 for annite and from 0.76 to 0.96 for siderophyllite. The H contents of the biotite do not vary with Fe^{3+} content for either composition and result in 94 and 100% of the OH sites being occupied in annite and siderophyllite, respectively. A consistent decrease in total octahedral cation content with increasing Fe^{3+} content for both annite and siderophyllite suggests that the Fe-vacancy substitution $3\text{Fe}^{2+} \rightleftharpoons 2\text{Fe}^{3+} + {}^{16}\text{Vac}$ plays a dominant role in the oxidation of these synthetic biotite samples.

INTRODUCTION

Biotite is among the most common rock-forming minerals and is found in assemblages of diverse bulk compositions and origins. Its chemical complexity is manifested by a wide range of both cationic and anionic substitutions (e.g., Deer et al., 1962; Guidotti, 1984). Such variability in biotite chemistry accommodates a number of possible mechanisms that can account for any observed compositional change. Even for the limited compositions of ideal annite, $\text{KFe}_3\text{AlSi}_3\text{O}_{10}(\text{OH})_2$, and siderophyllite, $\text{KFe}_{2.5}\text{Al}_2\text{Si}_{2.5}\text{O}_{10}(\text{OH})_2$, where Fe may occur in both the Fe^{2+} and Fe^{3+} states, many substitutions have been proposed to explain the oxidation of the Fe. Among the suggested substitutions are the following:



Two factors make it difficult to evaluate the relative roles of the different substitutions on the basis of available analytical data. First, electron microprobe analyses do not determine either the proportions of Fe^{3+} and Fe^{2+} or the concentrations of H. Second, the chemical complexities

of natural biotite make the unique determination of a set of substitutions difficult or impossible (Hewitt and Abrecht, 1986). Efforts to understand the Fe^{3+} substitution in biotite that circumvent one or the other or both of these issues have resulted in limited success.

Studies by Eugster and Wones (1962), Wones (1963a, 1963b), and Wones and Eugster (1965) suggested that the amount of Fe^{3+} in ferruginous biotite may be an indicator of the temperature and f_{H_2} under which the biotite crystallized. They proposed that oxidation occurs by the oxybiotite substitution. Wones and Eugster (1965) showed that the amount of Fe^{3+} increases with f_{O_2} at constant $f_{\text{H}_2\text{O}}$. Other studies using the Mössbauer effect for both synthetic (Wones et al., 1971; Ferrow and Annersten, 1984) and natural micas (Dyar, 1990; Guidotti and Dyar, 1991) have correlated an increase in Fe^{3+} with f_{O_2} .

The thermal oxidation of biotite in air and vacuum has been fairly well characterized by many workers (e.g., Vedder and Wilkins, 1969; Rimsaite, 1970; Hogg and Meads, 1975; Takeda and Ross, 1975; Bagin et al., 1980; Sanz et al., 1983; Ferrow, 1987a; Güttler et al., 1989; Tikhomi-rova et al., 1989) using a range of analytical tools including infrared spectroscopy, X-ray diffraction, differential thermal analysis, thermogravimetry, Mössbauer spectroscopy, EXAFS and XANES spectroscopies, and gas chromatography. From these studies there is a general consensus that biotite oxidizes by the oxybiotite substitution. Heating in vacuo does not change the mechanism, but it may delay oxidation until higher temperatures are reached (Sanz et al., 1983).

The substitution mechanisms for biotite oxidized by

* Present address: Department of Mineral Sciences, American Museum of Natural History, New York, New York 10024, U.S.A.

TABLE 1. Experimental conditions and wet-chemistry results

Expt.*	P (bars)	T (°C)	f_{H_2} (bars)	$\frac{Fe^{2+}}{Fe^{2+} + Fe^{3+}}$	H ₂ O (wt%)
Annite					
CR-41	1000	750	102.4	0.892	
DW-A34	1000	791	101.6	0.894	
DAH-92E4	1000	740	52.8	—	3.2–3.3
EP-182;282;7a82	1000	750	51.9	0.890	
DH-8-80	1000	750	51.9	0.886	
CR-10/11	1000	749	51.7	0.890	
CR-19/20	1000	640	51.7	0.900	
EP-14a83	1000	710	26.1	0.897	
EP-14b;15a	1000	700	26.1	0.844	
DAH-92E9	1000	602	10.0	—	3.3
CR-15/16	1000	640	10.0	0.828	
CR-23/24	1000	550	10.0	0.830	
EP-11a82;12a82;14a82;15a82	1000	615	6.0	0.814	
EP-2383 (NNO)	2000	600	4.0	0.769	
DAH-92E8	1000	519	1.0	—	3.3
CR-31;39	1000	520	1.0	0.765	
DW-A30 (HM)	2000	399	0.004	0.718	
Siderophyllite					
CR-17 (CCH ₄)	1000	801	143.0	0.954	
DAH-93E12	830	741	104.5	—	3.7
CR-26/30	1000	748	102.4	0.955	
DAH-93E6	1000	742	52.8	—	3.7
EP-3283	1000	750	51.9	0.971	
CR-6ACDE	1000	750	51.7	0.945	
CR-7BC	1000	669	51.7	0.940	
CR-13/22	1000	549	51.7	0.950	
EP-3083;21a83	1000	700	26.1	0.950	
DAH-93E7	1000	636	10.0	—	3.7
CR-3A/7A	1000	671	10.0	0.916	
CR-28/29	1000	581	10.0	0.913	
EP-3183;11a83	1000	615	6.0	0.895	
EP-2583 (NNO)	2000	600	4.0	0.829	
DAH-93E10	1000	517	1.0	—	3.7
CR-35/36	1000	520	1.0	0.861	
EP-1783 (HM)	2000	400	0.004	0.765	

* Data are averaged for experiments separated by semicolons. Solid buffer experiments are indicated: CCH₄ = graphite + methane; NNO = nickel + nickel oxide; HM = hematite + magnetite.

various chemical treatments have also been determined. Farmer et al. (1971) vermiculitized natural biotite, which resulted in a small degree of oxidation. Treatment with N₂H₄ reversed the changes completely. This result suggested that oxidation was accompanied by the release of protons, in accord with Reaction 2. More vigorous oxidation of the biotite by either H₂O₂ or a Br-saturated solution was not reversible and resulted in the formation of octahedral vacancies (Reaction 1). Similar experiments were carried out by Gilkes et al. (1972), who also found the dominant mechanism of oxidation to be the depletion of ⁶⁵Fe. This was accompanied by a small loss in inter-layer atoms (Reaction 4).

All the above studies involved biotite subjected to conditions well outside its stability field. Ferrow and Annersten (1984), Redhammer et al. (1993), and Rancourt et al. (1994) investigated synthetic ferruginous biotite that had been either synthesized or annealed using the solid O₂ buffer technique. All three studies used Mössbauer spectroscopy to characterize the oxidation state of the micas, and none of them determined the H₂O content of the biotite samples. Ferrow and Annersten (1984) con-

cluded that the oxybiotite substitution was the operative mechanism. Analyses of infrared spectra in the other two studies resulted in conflicting interpretations for the oxidation of Fe in annite: Redhammer et al. (1993) favored the ⁶⁵Fe vacancy substitution, and Rancourt et al. (1994) the oxybiotite substitution.

Geometric considerations of the mica structure impose further constraints on the amount of Fe³⁺ in biotite. In an ideal end-member annite, with all the Fe in the Fe²⁺ state, the tetrahedral sheet is smaller than the octahedral sheet. Substitution of cations smaller than Fe²⁺, such as Al³⁺, Mg²⁺, or Fe³⁺, into the octahedral sheet or substitution of cations larger than Si⁴⁺ into the tetrahedral sheet tends to stabilize the structure. Thus, Hazen and Wones (1972, 1978) predicted that Fe³⁺ occupies approximately 12 mol% of the octahedral sites or 7 mol% of the tetrahedral sites in annite. Studies by Wones et al. (1971), Partin (1984), Rebbert and Hewitt (1986), and Redhammer et al. (1993) have supported this prediction.

The purpose of this study is to investigate the relationships among f_{H_2} , structural H, and the oxidation state of Fe in synthetic biotite under equilibrium conditions.

EXPERIMENTAL METHODS AND ANALYTICAL TECHNIQUES

Biotite was synthesized from oxide mixes that had been reduced in a H_2 atmosphere at 500 °C for 1 h. Standard hydrothermal techniques were used for synthesis. The samples were encapsulated with excess distilled H_2O in either $Ag_{70}Pd_{30}$ or Au and held at a pressure of 4 kbar and temperatures between 400 and 600 °C in cold-seal pressure vessels (Tuttle, 1949). After 3–6 d the experimental products were examined optically with the polarizing microscope, and those with at least 98% biotite were utilized. Impurities identified include magnetite, fayalite, and sanidine. Most biotite specimens were then sealed in $Ag_{70}Pd_{30}$ capsules and annealed using Shaw vessel techniques for controlling f_{H_2} (Shaw, 1967; Hewitt, 1977). A few experiments were carried out using the solid O_2 buffering methods developed by Eugster (1957). Temperatures were measured with chromel-alumel thermocouples calibrated against the melting points of LiCl and KCl and found accurate to within ± 3 °C. Pressures were measured with Heise Bourdon-type gauges. One experiment was made with the graphite + CH_4 buffer (French and Eugster, 1965; Jacobs and Kerrick, 1981).

The annealed experiment products were examined optically, with most yields estimated by comparison with standard mixtures of synthetic fayalite in synthetic phlogopite (see Appendix A of Rebbert, 1986). Weight percent Fe^{2+} was determined by the wet-chemical technique outlined by Whipple (1974), and the stoichiometry of the biotite was checked by an ARL-SEM-Q microprobe. Because of the fine-grained nature of the crystals, the analytical methods and graphical procedures developed by Solberg et al. (1981) were employed. Only experiments where the percentage of biotite in the final product was $>93\%$ are reported in Table 1.

H contents of several synthetic annite and siderophyllite samples were determined by a procedure developed by Hewitt (in preparation). Dried samples of biotite weighing 5–10 mg are melted at 900 °C with an anhydrous $Na_2B_4O_7$ flux in open Pt capsules that are sealed together with Zn in evacuated SiO_2 glass tubes. The H_2 and H_2O released during melting is totally reduced to H_2 by the Zn in the system. The SiO_2 glass tube is then broken in an evacuated constant volume system, and the resulting pressure generated by the released H_2 is measured. Blanks are analyzed with every batch of samples, and the system is calibrated with a talc standard. Analysis of the standard talc by thermal decomposition yielded 4.76(3) wt% H_2O . Interlaboratory comparisons with D. Virgo and R. Popp at the Geophysical Laboratory showed excellent agreement on the H_2O contents of this standard material. The accuracy of the method is better than $\pm 2\%$ of the amount of H_2O present for the sample sizes and H_2O contents of the biotite analyzed in this study.

The X-ray diffraction patterns of the biotite samples were obtained from an automated Philips diffractometer

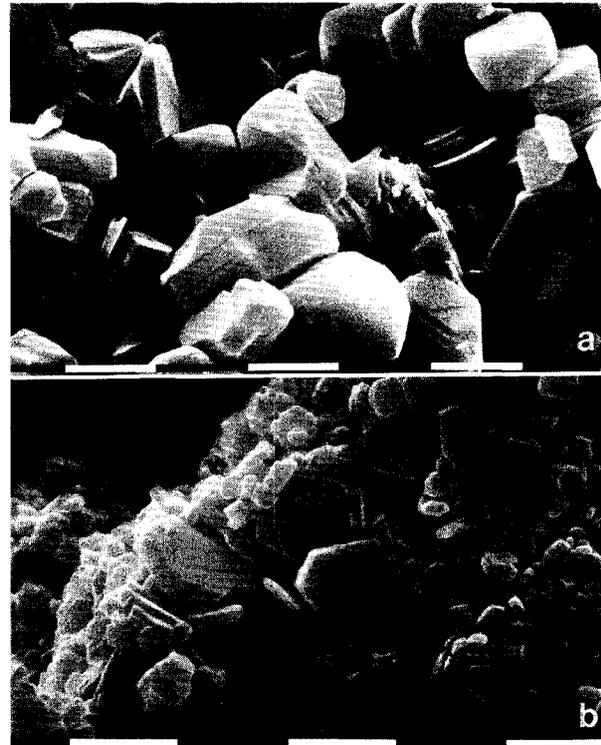


Fig. 1. SEM photomicrographs. (a) CR-10/11. (b) CR-35/36. Scale bar represents 1 μm .

using $CuK\alpha$ radiation. Fluorite, CaF_2 ($a = 5.4633 \pm 0.0005$ Å), served as an internal standard. Unit-cell dimensions were obtained by the refinement of 10–20 reflections for each sample (Appleman and Evans, 1973).

RESULTS

Optical examination

All the annealed experiment products were examined optically, and some were observed by SEM (Fig. 1). The biotite specimens are in general well-crystallized, with clean, smooth edges. Each experiment includes a wide range of crystal sizes, but low-temperature experiments have a smaller average size than higher temperature experiments, and siderophyllite specimens form smaller, thinner platelets (0.35–2 μm wide, <0.5 μm thick) than the annite specimens (0.5–2 μm , 0.25–1 μm thick) under similar high-temperature conditions. The siderophyllite is a bluish green color, whereas the annite is a darker olive drab, and both show a faint pink to green normal pleochroism. All experiments include a small percentage of sanidine and magnetite or fayalite (or both) as impurity phases. They constitute 2–5% (by volume) of the siderophyllite samples and 4–7% of the annite samples.

Chemical analysis

The wet-chemical analyses show a positive correlation between the percentage of Fe^{2+} and f_{H_2} (Table 1, Fig. 2).

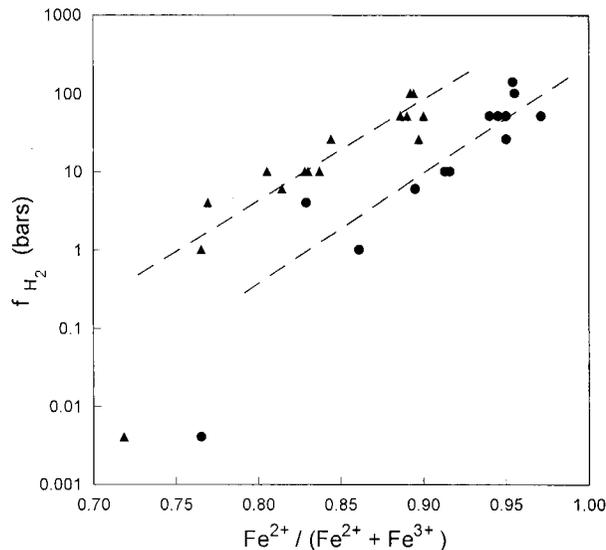


Fig. 2. Hydrogen fugacity vs. Fe^{2+} content for the synthetic biotite samples in Table 1. Triangles represent annite data, circles the siderophyllite data. Errors related to the Fe^{2+} determination are within the size of the symbols.

The annite contains 5–9 mol% more Fe^{3+} than siderophyllite annealed under similar conditions. The data in Table 1 indicate that the amount of Fe^{3+} is constant [± 0.001 – $0.03 \text{ Fe}^{2+}/(\text{Fe}^{2+} + \text{Fe}^{3+})$] for a given f_{H_2} regardless of temperature variations of up to 200 °C. This agrees well with the results of Wones (1963b), who showed that biotite samples synthesized at the same temperature and total pressure and similar f_{H_2} , but at different values of f_{O_2} as a result of a different C-O-H fluid, had essentially identical γ refractive indices. The refractive index was shown to be inversely related to f_{H_2} and therefore to Fe^{2+} content.

Electron microanalysis of some of the biotite specimens indicates that their stoichiometry agrees well with that of the two ideal end-member compositions (Table 2). The annite specimens tend to be slightly deficient in

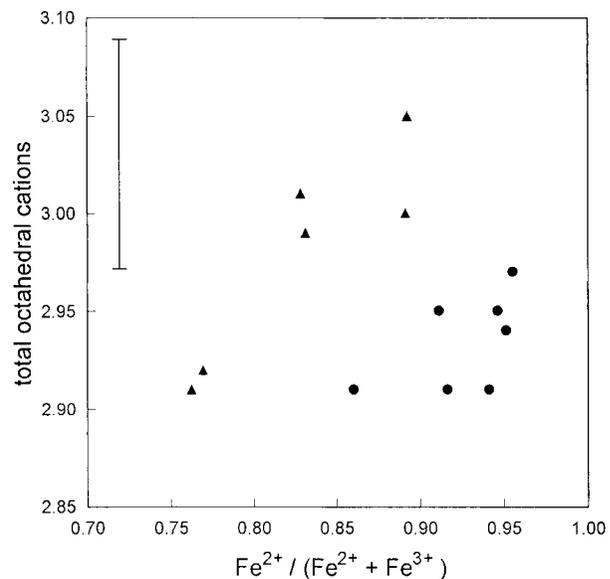


Fig. 3. Octahedral occupancies vs. $X_{\text{Fe}^{2+}}$. Symbols as in Fig. 2; error bar represents 2σ .

SiO_2 and H_2O , and both compositions have about 10% interlayer vacancies. All the siderophyllite samples have octahedral vacancies, whereas both octahedral vacancies and filled sites occur in annite. There is a general trend of decreasing total octahedral cation content with increasing Fe^{3+} that is especially apparent for the annite experiments (Fig. 3).

Foster (1960) assumed that $^{49}\text{Fe}^{3+}$ is present in trioctahedral mica compositions with insufficient Si and Al to fill the site. Since that time, Annersten (1975), Dyar and Burns (1986), Dyar (1987), and Rancourt et al. (1992) have verified its existence by means of Mössbauer spectroscopy. On the basis of the calculated cation occupancies, we expect to see $^{49}\text{Fe}^{3+}$ in annite, since the sum of Si and Al cations in all cases is < 4 . Siderophyllite, on the other hand, has sufficient Si and Al to completely fill the tetrahedral site.

TABLE 2. Chemical composition of synthetic biotite as weight percent oxides and numbers of cations based on 12 anions

Sample*	41	10/11	15/16	23/24	31	39	17	26/30	6ACDE	7BC	13/22	3A/7A	28/29	35/36
SiO_2	33.6	34.2	32.9	33.8	34.0	34.1	30.3	30.9	31.2	31.8	30.6	31.4	30.6	30.2
Al_2O_3	10.6	10.6	10.1	9.9	9.8	10.2	21.0	20.5	20.7	20.5	20.6	20.8	20.6	21.1
Fe_2O_3	5.3	5.2	8.5	8.3	11.2	11.4	1.8	1.8	2.2	2.3	2.0	3.3	3.5	5.5
FeO	39.2	38.2	36.8	36.6	33.6	32.8	34.3	34.4	33.6	32.8	33.9	32.0	33.0	30.8
K_2O	8.1	8.5	8.4	8.1	8.1	8.2	8.9	8.6	8.5	8.7	9.2	8.8	8.6	8.7
H_2O	3.3	3.3	3.3	3.3	3.3	3.3	3.7	3.7	3.7	3.7	3.7	3.7	3.7	3.7
Total	100.1	100.0	100.0	100.0	100.0	100.0	100.0	99.9	99.9	99.8	100.0	100.0	100.0	100.0
Si	2.86	2.90	2.81	2.87	2.87	2.87	2.48	2.53	2.54	2.58	2.51	2.55	2.50	2.46
Al	1.06	1.06	1.02	0.99	0.97	1.01	2.03	1.98	1.99	1.96	1.99	1.99	1.98	2.02
Fe^{3+}	0.34	0.33	0.55	0.53	0.71	0.72	0.11	0.11	0.13	0.14	0.12	0.20	0.22	0.34
Fe^{2+}	2.79	2.71	2.63	2.60	2.37	2.31	2.35	2.35	2.29	2.23	2.32	2.17	2.25	2.09
Σ_{oct}	3.05	3.00	3.01	2.99	2.92	2.91	2.97	2.97	2.95	2.91	2.94	2.91	2.95	2.91
K	0.88	0.92	0.92	0.88	0.87	0.88	0.93	0.90	0.88	0.90	0.96	0.91	0.90	0.90
OH	1.88	1.87	1.88	1.87	1.86	1.85	2.02	2.02	2.01	2.00	2.02	2.00	2.01	2.01
O	10.12	10.13	10.12	10.13	10.14	10.15	9.98	9.98	9.99	10.00	9.98	10.00	9.99	9.99

* All sample numbers preceded by CR-.

The H₂O contents of biotite synthesized and annealed under conditions identical to those for the determination of Fe³⁺ contents are all very consistent (Table 1). The annite samples average 3.3(1) wt% H₂O and show no regular variation with Fe³⁺ contents. Similarly, the siderophyllite samples average 3.70(5) wt% H₂O and also show no regular variation with Fe³⁺ contents. By comparison with the ideal of 3.52 wt% H₂O in annite and 3.63 wt% H₂O in siderophyllite, it is apparent that there is a small deficiency (~6%) in the annite and a stoichiometric amount or possibly a slight excess (~1%) of H₂O in the siderophyllite. What is most important for understanding the Fe³⁺ substitution in biotite is that the data demonstrate no correlation between the Fe³⁺ and the H contents of these synthetic biotite specimens.

X-ray data

Results from the X-ray data refinements are shown in Figure 4. There are good correlations between the mole fraction of Fe²⁺ and the *b*-cell dimension, the *c*-cell dimension, and the unit-cell volume. Other workers demonstrated similar trends (Eugster and Wones, 1962; Wones, 1963b; Rutherford, 1973; Ohta et al., 1982; Ferron and Annersten, 1984). Gilkes et al. (1972) demonstrated that *c* remained constant with increasing oxidation of mica. However, their starting mica is fairly Mg-rich, and Wones (1963b) showed that the progressive change in cell dimensions with oxidation is much less for magnesian biotite. The absolute values for the cell edges and unit-cell volumes of our biotite specimens are similar to those from the above studies, excluding the natural biotite of Ohta et al. (1982), which has a very different composition.

Errors in Fe²⁺ and Fe³⁺ determinations

The errors introduced in calculating the mole fraction of Fe²⁺ are difficult to determine. The precision of the titrations is quite good, with replicate determinations agreeing to within 0.1–0.3% (relative). If all impurities present are taken to be Fe-bearing phases, then absolute maximum errors can be calculated. The presence of 3 wt% fayalite leads to overestimation of the Fe²⁺ content of the biotite, but because the weight percent FeO is similar in fayalite and iron biotite, this amounts to a maximum correction of 0.2–0.4 mol% of Fe²⁺. If instead there were 3 wt% of magnetite present, the Fe²⁺ content would be underestimated by a maximum of 2.5 mol%. The presence of sanidine decreases the correction, since it suggests that the sum of the impurities has a stoichiometry equivalent to that of the biotite. On the basis of the optical examination of these experiments, magnetite constitutes at most 2%, but usually <1%, of the total weight of material. This corresponds to underestimating the Fe²⁺ content by a maximum of 1 mol%.

Errors based on the linear regression statistics from the microprobe analyses were also determined for CR-17. The sample variance of the mean elemental weight percent for a given total weight percent was calculated. Weight

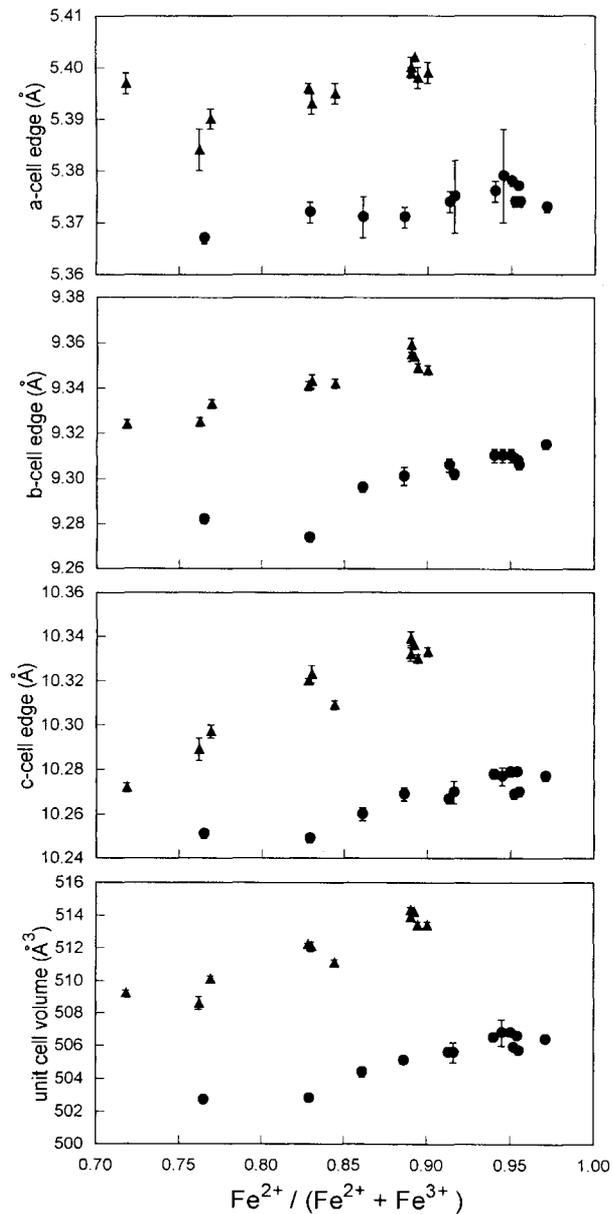


Fig. 4. X-ray data vs. $X_{Fe^{2+}}$. Symbols as in Fig. 2.

percent FeO was adjusted to account for the presence of Fe₂O₃, and the errors were propagated through this calculation. The standard deviation for each oxide was used to generate a random set of 100 oxide weight percents with a normal distribution about the mean. These in turn were used to calculate mineral formulas whose mean and standard deviation were determined. The error on Fe²⁺ is ±0.05 cations and on Fe³⁺ is ±0.02 cations.

DISCUSSION

The variation of Fe²⁺ with f_{H_2}

The Fe²⁺/(Fe²⁺ + Fe³⁺) ratios of annite and siderophyllite specimens as a function of f_{H_2} are shown in Fig-

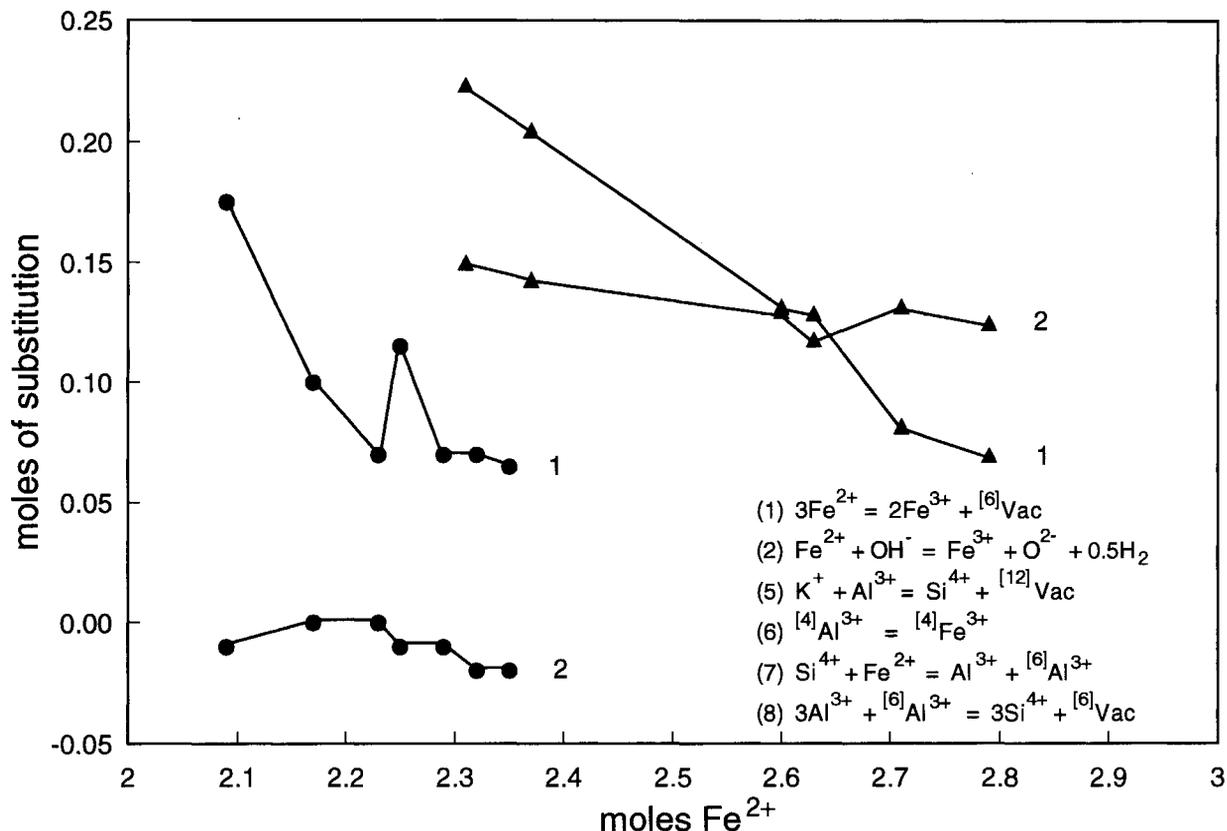


Fig. 5. Moles of substitutions 1 and 2 required to account for differences in composition between the two ideal end-members and the synthetic biotite samples as a function of total number of Fe²⁺. Symbols as in Fig. 2. The other four substitutions (5–8) used to describe the synthetic biotite compositions are listed as well.

ure 2. For annite, the maximum attainable Fe²⁺ content is 0.90 mole fraction. This value is reached at a f_{H_2} of 50 bars and does not increase at f_{H_2} values up to 102 bars. There is a regular decrease in the amount of Fe²⁺ at lower f_{H_2} , with the lowest value of $X_{\text{Fe}^{2+}} = 0.72$ at the HM buffer. It should be noted that the data are not isothermal, with temperatures ranging from approximately 800 °C at the highest f_{H_2} values to 400 °C at the HM buffer. Most experiments were carried out near the upper thermal stability for biotite at the various oxidation conditions. However, there are numerous examples in Table 1 (e.g., compare CR-10/11 and CR-19/20, CR-15/16 and CR-23/24, and CR-6ACDE, CR-7BC and CR-13/22), demonstrating that the Fe²⁺ content does not change with temperature, and therefore f_{O_2} , under conditions of constant f_{H_2} . Significant differences in Fe²⁺/(Fe²⁺ + Fe³⁺) ratios are generated only by differences in bulk composition and differences in f_{H_2} .

The data for annite support the predictions of Hazen and Wones (1972, 1978) with regard to the structural limitation on the maximum amount of Fe²⁺ allowable for a fit between the tetrahedral and octahedral sheets. In addition, they clearly demonstrate the existence of a significant variation in Fe²⁺ content with f_{H_2} . The trend for

siderophyllite is similar, except that the maximum Fe²⁺ content is approximately 0.96 mole fraction Fe²⁺/(Fe²⁺ + Fe³⁺). Because Tschermak's substitution (Fig. 5, Reaction 7) increases the lateral dimension of the tetrahedral sheet and decreases that of the octahedral sheet, this higher limit for the Fe²⁺ content provides further support for the Hazen and Wones's (1972, 1978) arguments. Simple calculations on the size changes for the tetrahedral and octahedral sheets with Tschermak's substitution indicate that the siderophyllite composition should have no required Fe³⁺ caused by size limitations. Whether the 0.04 mole fraction Fe³⁺ in the siderophyllite samples synthesized at very reducing conditions is present because of a lack of analytical accuracy, or whether it represents ¹⁴⁹Fe³⁺, or whether it is simply the result of not having a sufficiently high f_{H_2} is not known. However, the consistently higher Fe²⁺ content of siderophyllite is a strong argument for the Hazen and Wones (1972, 1978) model.

Within the limits of the data, the slopes of the variation of Fe²⁺/(Fe²⁺ + Fe³⁺) vs. f_{H_2} are similar, if not identical, for annite and siderophyllite (Fig. 2). This suggests that the same oxidation-substitution mechanism is active for both compositions. Although there is no a priori reason that the variation be linear, and clearly it cannot be so at

gested two main substitutions to accommodate trivalent octahedral cations: (a) $Al^{3+} + R^{3+} \rightleftharpoons Si^{4+} + R^{2+}$ and (b) $3R^{2+} \rightleftharpoons 2R^{3+} + {}^{69}Vac$, which corresponds to the Fe vacancy reaction. Most natural biotite specimens require both substitutions to be operative, but Foster (1960) demonstrated that the relative importance of Reaction a diminishes with increasing Fe content. The synthetic mica data support that conclusion because when they are fitted to a different set of substitutions that includes Foster's Reaction a, they show only a weak correlation between the amount of Reaction a required and the degree of oxidation. Furthermore, the Fe vacancy substitution is consistent with the results and conclusions of both the hydrothermal study of Redhammer et al. (1993) and the two studies that treated biotite with chemical oxidants, Farmer et al. (1971) and Gilkes et al. (1972).

Finally, Hazen and Wones (1972) and Ferrow (1987b) found that while octahedral cation substitutions have a dramatic effect on the *b*-cell dimension, they have virtually no effect on *c*. The Fe vacancy substitution involves only octahedral sites, and so it is expected that *c* will remain constant. Yet Wones (1963b), Rutherford (1973), Ferrow and Annersten (1984), Redhammer et al. (1993), and our data all demonstrate a decrease in *c* with hydrothermal oxidation of biotite (Fig. 4). A possible explanation for this observation lies in the orientation of the OH dipole. Sanz et al. (1984) demonstrated that OH bonds near octahedral vacancies are inclined away from the *c* axis. This inclination diminishes the repulsion between the positive K and H ions and is used by Gilkes et al. (1972) to explain the increased stability of K in oxidized biotites. This should result in diminishing the K-O bond length and thereby decreasing the *c*-cell dimension.

The Fe vacancy reaction does not directly explain the pronounced correlation between f_{H_2} and the mole fraction of Fe^{2+} (Fig. 2), and so far we have not explained why the results from our hydrothermal experiments are so disparate from the numerous thermal oxidation experiments that support an oxybiotite substitution. This behavior is irreconcilable if only the biotite is considered. However, the hydrothermal experiments involve the system biotite- H_2O , and the thermal oxidation experiments take place in the system biotite-air. From an energetic point of view, there is probably little alternative to an oxybiotite substitution when biotite is heated under disequilibrium conditions in the presence of air (~ 0.2 bars f_{O_2}) or even in the presence of a modest vacuum, where the partial pressure of O_2 might be on the order of 10^{-7} bars. In the hydrothermal system biotite- H_2O , a variety of oxidation reactions involving the two phases could easily occur. Even more are possible if we expand the system to include the capsule or other solid phases. One simple reaction for the biotite- H_2O system that is consistent with both the Fe vacancy substitution and the inverse dependence between Fe^{3+} in biotite and f_{H_2} is the following: $3Fe_{(o)}^{2+} + H_2O_{(f)} \rightleftharpoons 2Fe_{(o)}^{3+} + Fe_{(f)}^{2+} + O_{(f)}^{2-} + H_{2(f)}$. Here, a decrease in f_{H_2} drives the reaction to the right and thus to more Fe^{3+} -rich compositions in the biotite. Our data

are insufficient to verify its occurrence, but they are at least consistent with a reaction of this sort.

We propose that the mechanism of oxidation in Fe-bearing silicates is linked both to environmental conditions and to crystal chemical considerations. Certain natural biotite specimens have been shown to oxidize, at least in part, by the oxybiotite substitution. Takeda and Ross (1975) described the occurrence of biotite from the top of a series of rhyodacite flows in which 53% of the OH deficiency can be attributed to the oxybiotite reaction in a highly oxidized and probably nearly anhydrous environment. Feldstein et al. (1993) correlated Fe^{3+} with H contents in phlogopite from a variety of mantle-derived melts, but these biotites represent compositions and have formed under conditions far different from those for our experiments. The current data supporting the vacancy substitution are for iron biotite synthesized under conditions of high f_{H_2O} (400–1000 bars).

Amphiboles have long been assumed to oxidize by the equivalent oxyamphibole substitution (e.g., Clowe et al., 1988). Investigations by Popp et al. (1993) and Phillips et al. (1993) have demonstrated clearly that this assumption is well founded for experiments carried out under both hydrothermal and atmospheric conditions. The significant difference in oxidation behavior between biotite and amphibole must be related to the common tendency for octahedral vacancy formation in the micas compared with the rare occurrence of substitutions in the amphiboles that generate octahedral vacancies.

The jump from understanding the Fe^{3+} - Fe^{2+} substitutions in synthetic iron biotite to dealing with the complex compositions of natural biotite is a significant one. The following observations are based on the extensive amount of analytical data in the literature for the compositions of natural biotite: (1) Fe^{3+} - Fe^{2+} ratios and total Fe contents vary widely. (2) Octahedral vacancies are common and variable in amount. (3) The OH contents are often less than the ideal amount. (4) Ti is frequently present as a minor or major element. (5) K and Na contents are commonly insufficient to fill the interlayer site. (6) The Si-Al ratio is highly variable. The Fe^{3+} variation that is observed in natural and synthetic biotite can legitimately be related to substitutions involving any of the constituents in the final five observations. That we have been able to eliminate the oxybiotite substitution and demonstrate the importance of the ${}^{69}Fe$ vacancy substitution to account for the variation of Fe^{3+} in some simple synthetic systems is significant. To generalize very far from that success is not well justified. Similarly, to find a correlation between Fe^{3+} and OH content in a suite of natural biotite samples is interesting but does not establish the oxybiotite substitution as generally important and, depending on the other compositional variations in the same samples, might not even establish it as important in the samples described. It is apparent that the Fe^{3+} substitutions in biotite depend on the composition of the biotite as well as the intensive parameters of the environment. We suggest that additional experimental and crystal-

chemical studies on synthetic or generally less compositionally complex natural biotite specimens must be undertaken to understand these relationships further.

ACKNOWLEDGMENTS

This material is based in part upon work supported under a National Science Foundation graduate fellowship to C.R.R. We are grateful to everyone who gave support and advice during the course of this project, especially D.R. Wones and J. Abrecht. We appreciate the sharp SEM photographs taken by J. Abrecht and the assistance of T.N. Solberg with the microprobe analyses. We thank Darby Dyar and an anonymous reviewer for helpful reviews. C.R.R. thanks Dave Wayne for a helpful review of an early version of the manuscript.

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MANUSCRIPT RECEIVED JUNE 27, 1994

MANUSCRIPT ACCEPTED DECEMBER 9, 1994