TABLE 1

Author	Specimen locality	Method and results of analysis	Proposed formulae
Van Tassel (1959)	Argenteau, Belgium	Chemical Analysis (recalculated) Al: 18.8 F: 16.0 PO_4: 27.6 OH: 6.5 } OH calculated value to balance Al charge. H_2O: 31.1 } H_2O by difference. D_m = 2.12. Powder data for material from Cornwall, Bavaria, and Belgium.	$3Al_2O_3 \cdot 4AlF_3 \cdot 2P_2O_5 \cdot 27H_2O$ (from Van Tassel, no OH content) $Al_{2.4}PO_4F_{2.9}(OH)_{1.3} \cdot 6H_2O$ (calculated, assuming Al balanced by OH)
Chukrov (1963)	Kazakhastan	Chemical Analysis: Al: 16.83 17.04 F: 14.80 14.62 PO_4: 28.44 28.97 OH: 3.32 3.59 } OH calculated to H_2O: 35.89 35.36 } balance Al charge. etc: 0.77 0.33 D_m = 2.17 (both specimens) Powder data for Kazakhastan material.	$Al_2(F \cdot OH)_3PO_4 \cdot 7H_2O$ (may be 6 or 7 H_2O per formula unit. F:OH is 4:1)

GUY, B. B. AND G. A. JEFFREY (1966) The Crystal Structure of Fluellite, $Al_2PO_4F_2(OH) \cdot 7H_2O$. *Amer. Mineral.* **51,** 1579–1592.

VAN TASSEL, R. (1959) Autunite, apatite, delvauxite, évansite et fluellite de la région de Visé. *Bull. Soc. belge Géol.*, **68,** 226–248.

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BARIUM-VANADIUM MUSCOVITE AND VANADIUM TOURMALINE FROM MARIPOSA COUNTY, CALIFORNIA: A CORRECTION

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Dr. L. G. Berry has drawn my attention to the absence of the (006) and (024) reflections from the X-ray data for barium-vanadium muscovite in my paper (*Amer. Mineral.* **51,** 1623–1639); both reflections are typical of $2M_1$ muscovite. I neglected to include (006) in the data; it has d (meas.) = 3.339 Å, and an intensity of 90. These values were obtained from film because in diffractometer work the 2θ range in question was swamped by the intense internal standard quartz peak; d (calc.) of (006) is 3.336 Å. I find no indication of (024) on the X-ray film of the mica; (024) is considerably less intense than nearby (006) (*cf.* ASTM card 6-0263), and presumably is masked by (006).

ERRATA

NESTER, J. F. (1967) Growth of synthetic calcite crystals. **52,** 276–280: p. 276, for "Li_2O_3" read "La_2O_3."