very useful while doing geologic mapping in remote areas. During mapping, the staining has been carried out on individual specimens during traverses or, more commonly, several specimens were stained at the end of a day at a base camp.

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THE CRYSTAL STRUCTURE OF FLUELLITE: A CORRECTION

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In a recent X-ray single-crystal study of a specimen of the mineral fluellite, $Al_2PO_4F_2OH \cdot 7H_2O$, from Cornwall, England, Guy and Jeffrey (1966) determined the presence of the phosphate ion in the structure, although some previous investigations had suggested the composition was that of an aluminum hydroxy-fluoride monohydrate. Two publications, one by Van Tassel (1959) and the other published by the USSR Academy of Science (1963), also indicating the presence of phosphate in this mineral, have been noted by one of us (RVT).

Since reference to these publications was overlooked in the crystal structure paper, a summary of the chemical composition data from them is presented in Table 1.

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TABLE 1

Author	Specimen locality	Method and results of analysis	Proposed formulae
Van Tassel (1959)	Argenteau, Belgium	Chemical Analysis (recalculated) Al: 18.8 F: 16.0 PO_4: 27.6 OH: 6.5 } OH calculated value to balance Al charge. H_2O: 31.1 } H_2O by difference. $D_m = 2.12$. Powder data for material from Cornwall, Bavaria, and Belgium.	$3Al_2O_3 \cdot 4AlF_3 \cdot 2P_2O_5 \cdot 27H_2O$ (from Van Tassel, no OH content) $Al_{2.4}PO_4F_{2.9}(OH)_{1.3} \cdot 6H_2O$ (calculated, assuming Al balanced by OH)
Chukrov (1963)	Kazakhastan	Chemical Analysis: Al: 16.83 17.04 F: 14.80 14.62 PO_4: 28.44 28.97 OH: 3.32 3.59 } OH calculated to H_2O: 35.89 35.36 } balance Al charge. etc: 0.77 0.33 $D_m = 2.17$ (both specimens) Powder data for Kazakhastan material.	$Al_2(F \cdot OH)_3PO_4 \cdot 7H_2O$ (may be 6 or 7 H_2O per formula unit. F:OH is 4:1)

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BARIUM-VANADIUM MUSCOVITE AND VANADIUM TOURMALINE FROM MARIPOSA COUNTY, CALIFORNIA: A CORRECTION

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Dr. L. G. Berry has drawn my attention to the absence of the (006) and (024) reflections from the X-ray data for barium-vanadium muscovite in my paper (*Amer. Mineral.* **51,** 1623–1639); both reflections are typical of $2M_1$ muscovite. I neglected to include (006) in the data; it has d (meas.) = 3.339 Å, and an intensity of 90. These values were obtained from film because in diffractometer work the 2θ range in question was swamped by the intense internal standard quartz peak; d (calc.) of (006) is 3.336 Å. I find no indication of (024) on the X-ray film of the mica; (024) is considerably less intense than nearby (006) (*cf.* ASTM card 6-0263), and presumably is masked by (006).

ERRATA

NESTER, J. F. (1967) Growth of synthetic calcite crystals. **52,** 276–280: p. 276, for "Li_2O_3" read "La_2O_3."