

## NEW MINERAL NAMES

### Stiepelmannite

PAUL RAMDOHR AND E. THILO: Stiepelmannite, ein neues Mineral der Haminitgruppe mit Yttrium und seltenen Erden. *Centr. Mineral.* Abt. A, Nr. 1, 1-8, 2 figs., 1940.

NAME: For Mr. Stiepelmann, owner and operator of the Klein Spitzkopje property, who furnished the material.

CHEMICAL PROPERTIES: A phosphate of aluminum and yttrium earths. Formula  $\text{YPO}_4 \cdot \text{AlPO}_4 \cdot 2\text{Al}(\text{OH})_3$ . Analysis:  $\text{SiO}_2$  0.29,  $\text{X}_2\text{O}_3$  29.25,  $\text{ZrO}_2$  1.12,  $\text{CaO}$  0.50,  $\text{Al}_2\text{O}_3$  30.83,  $\text{H}_2\text{O}$  11.07,  $\text{P}_2\text{O}_5$  (by difference) 26.94; Sum 100.00. Attacked with difficulty by concentrated hydrochloric acid.

CRYSTALLOGRAPHIC PROPERTIES: Trigonal. Habit rhombohedral, resembling modified cubes.  $c=1.272$ , Forms  $(10\bar{1}\bar{1})$ ,  $(0001)$ ,  $(02\bar{2}\bar{1})$ .  $a=6.75 \text{ \AA}$ ,  $c=16.52 \text{ \AA}$ .

PHYSICAL AND OPTICAL PROPERTIES: Colorless to pale wine yellow. Luster high.  $H=6$ . Fracture conchoidal. Cleavage  $(0001)$  and  $(11\bar{2}0)$  indistinct.  $G=3.671-3.713$ . Uniaxial, positive.  $\omega=1.695$ ,  $\epsilon=1.705$ .

OCCURRENCE: Found in small crystals, coated with hyalite and limonite upon albite, microcline, fluorite, etc., from the beryl pegmatites of Klein Spitzkopje, South West Africa.

REMARKS: Stiepelmannite is a member of the alunite-jarosite group and is closely related to florencite.

W. F. FOSHAG

### DISCREDITED SPECIES

#### Kreuzbergite (=fluellite)

A. SCHOLZ AND H. STRUNZ: Identität von Kreuzbergite mit Fluellite. *Centr. Mineral.* Abt. A, Nr. 6, 133-137 (1940).

Reexamination of kreuzbergite shows it is not a phosphate, but from its optical, crystallographical and Debye-Scherrer diagrams is fluellite  $\text{Al}(\text{F}, \text{OH})_3 \cdot \text{H}_2\text{O}$ .

	Stenna Gwyn	Hagendorf	Pleystein	Konigswart
$\alpha$	1.475	1.490	1.489	
$\beta$	1.490	1.496	1.495	
$\gamma$	1.510	1.509	1.506	
$a$	0.770	0.7724	0.7485	0.7516
$c$	1.874	1.8715	1.9008	1.7794

The variations are explained as due to isomorphous replacement of (F) by (OH.)

W. F. F.